Science 10 Chemical Reactions Review Name: _____ /40

1. Name the following molecular (covalent) compounds. (nm-nm and prefixes) (3 pts)

   a. SeO_2 selenium dioxide

   b. N_2Cl_4 dinitrogen tetrachloride

   c. Br_2F_5 dibromium pentafloride

2. Write the formulas for the following molecular (covalent) compounds. (3 pts)

   a. iodine monochloride ICl

   b. tetraphosphorous hexasulfide P_4S_6

   c. boron trichloride BCl_3

3. Write the formulas for the following **ionic** compounds. (mn-nm) (10 pts) Show work

Name	Formula
magnesium oxide	$Mg^{+2}O^{-2} \rightarrow MgO$
aluminum oxide	$Al^{+3}O^{-2} \rightarrow Al_2O_3$
manganese (IV) oxide	$Mn^{+4}O^{2-} \rightarrow Mn_2O_4 \rightarrow MnO_2$
mercury (II) sulphide	$Hg^{2+}S^{2-} \rightarrow Hg_2S_2 \rightarrow HgS$
chromium(III)chloride	$Cr^{3+}Cl^{-} \rightarrow CrCl_3$
cadmium nitride	$Cd^{2+}N^{-3} \rightarrow Cd_3N_2$
gallium nitrite	$Ga^{+3}(NO_2)^{-} \rightarrow Ga(NO_2)_3$
nickel (II) acetate or $Ni(CH_3COO)_2$	$Ni^{2+}C_2H_3O_2^{-} \rightarrow Ni(C_2H_3O_2)_2$
silver sulfate	$Ag^{+}SO_4^{-2} \rightarrow Ag_2SO_4$
platinum (IV) phosphate	$Pt^{4+}PO_4^{3-} \rightarrow Pt_3(PO_4)_4$